

Write the ionic formula

Name _____

Date _____

Balance the charges.

Problem	Answer	Problem	Answer
1. $\text{Na}^+ + \text{O}^{2-}$		6. $\text{Fe}^{3+} + \text{OH}^-$	
2. $\text{Mg}^{2+} + \text{Cl}^-$		7. $\text{NH}_4^+ + \text{CO}_3^{2-}$	
3. $\text{Al}^{3+} + \text{S}^{2-}$		8. $\text{Cu}^{2+} + \text{NO}_3^-$	
4. $\text{K}^+ + \text{SO}_4^{2-}$		9. $\text{Ba}^{2+} + \text{O}^{2-}$	
5. $\text{Ca}^{2+} + \text{PO}_4^{3-}$		10. $\text{Pb}^{4+} + \text{O}^{2-}$	

Answer key — Write the ionic formula

#	Problem	Answer
1	$\text{Na}^+ + \text{O}^{2-}$	Na_2O
2	$\text{Mg}^{2+} + \text{Cl}^-$	MgCl_2
3	$\text{Al}^{3+} + \text{S}^{2-}$	Al_2S_3
4	$\text{K}^+ + \text{SO}_4^{2-}$	K_2SO_4
5	$\text{Ca}^{2+} + \text{PO}_4^{3-}$	$\text{Ca}_3(\text{PO}_4)_2$
6	$\text{Fe}^{3+} + \text{OH}^-$	$\text{Fe}(\text{OH})_3$
7	$\text{NH}_4^+ + \text{CO}_3^{2-}$	$(\text{NH}_4)_2\text{CO}_3$
8	$\text{Cu}^{2+} + \text{NO}_3^-$	$\text{Cu}(\text{NO}_3)_2$
9	$\text{Ba}^{2+} + \text{O}^{2-}$	BaO
10	$\text{Pb}^{4+} + \text{O}^{2-}$	PbO_2